

Information sheet for the course Inorganic Chemistry of Materials

University: <i>Alexander Dubček University of Trenčín</i>						
Faculty: <i>Faculty of Industrial Technologies in Púchov</i>						
Course unit code: <i>MI-P-2</i>				Course unit title: <i>Inorganic Chemistry of Materials</i>		
Type of course unit: <i>compulsory</i>						
Planned types, learning activities and teaching methods: <i>Lecture: 2 hours weekly/26 hours per semester of study</i> <i>Seminar: 2 hours weekly/26 hours per semester of study</i> <i>Laboratory tutorial: 2 hours weekly/26 hours per semester of study</i>						
Number of credits: <i>6</i>						
Recommended semester: <i>1st semester in the 1st year full-time</i> <i>1st semester in the 1st year part-time</i>						
Degree of study: <i>the 1st degree of study (Bachelor's degree)</i>						
Course prerequisites: <i>none</i>						
Assessment methods: <i>Evaluation of course includes partial evaluation; basic characteristic of static of chemical substances, chemical reaction (thermodynamic and kinetic aspects), kind of chemical bond, physical properties of inorganic substances</i>						
Learning outcomes of the course unit: <i>Student profits the survey on the inorganic chemistry and materials.</i>						
Course contents: <ol style="list-style-type: none"> <i>1. States of chemical substances – basic characteristic</i> <i>2. Chemical thermodynamic: Enthalpy, free energy, Entropy, Gibbs energy</i> <i>3. Chemical equilibrium, equilibrium constant</i> <i>4. Kinetics: rate laws</i> <i>5. The effect of concentration, temperature and catalyst on reaction rate</i> <i>6. Acids and basis (Arrhenius, Bronsted and Lewis theory)</i> <i>7. Protolytic reactions: neutralization and hydrolysis</i> <i>8. Precipitation reactions: product of solubility</i> <i>9. Redox reactions: reducing and oxidizing agents, redox potential</i> <i>10. Reactions of complex formation: complex, chromophore</i> <i>11. Wave mechanics: wave function, atomic orbitals, electron configurations (the aufbau principle, Hund's rule, the Pauli principle)</i> <i>12. Physical essence of chemical bond</i> <i>13. Kinds of chemical bonds</i> <i>14. Electric, magnetic, optical and thermal properties of inorganic substances</i> 						
Recommended of required reading: <ol style="list-style-type: none"> <i>1. Jóna E., Ondrušová D., Pajtášová M.: Priemyselná anorganická chémia I: Všeobecná časť, 2007</i> <i>2. Garaj J.: Chémia učebné texty pre mechanické odbory, Trenčín,</i> 						
Language: <i>Slovak</i>						
Remarks: <i>none</i>						
Evaluation history:						
A	B	C	D	E	FX	

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Lecturers: <i>prof. Ing. Eugen Jóna, DrSc., prof. Ing. Darina Ondrušová, PhD.</i>					
Last modification: <i>31.03.2014</i>					
Supervisor: <i>prof. Ing. Darina Ondrušová, PhD.</i>					